

2-Nitro-*N*-(4-pyridinio)benzene-sulfonamide monohydrateLiang-Bin Hu,<sup>a\*</sup> Jian Chen<sup>b</sup> and Chang-Zhong Liu<sup>c</sup>

<sup>a</sup>School of Food Science, Henan Institute of Science and Technology, Xinxiang 453003, People's Republic of China, <sup>b</sup>College of Life Sciences, Nanjing Agricultural University, Nanjing 210095, People's Republic of China, and <sup>c</sup>College of Animal Science, Henan Institute of Science and Technology, Xinxiang 453003, People's Republic of China

Correspondence e-mail: liangbin\_hu@yahoo.cn

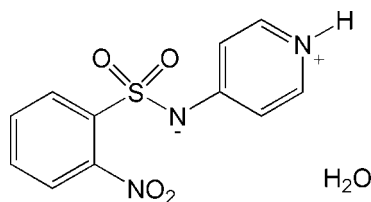
Received 22 October 2008; accepted 30 October 2008

Key indicators: single-crystal X-ray study;  $T = 113$  K; mean  $\sigma(\text{C}-\text{C}) = 0.004$  Å;  $R$  factor = 0.032;  $wR$  factor = 0.090; data-to-parameter ratio = 12.0.

The title compound,  $\text{C}_{11}\text{H}_9\text{N}_3\text{O}_4\text{S}\cdot\text{H}_2\text{O}$ , contains both an acid and a base centre and displays a zwitterionic structure. There are two independent molecules and two water molecules in the asymmetric unit. The dihedral angles between the benzene ring and the pyridinium ring are  $109.7$  (1) and  $110.7$  (1)°. The dihedral angles between the nitro group and the benzene ring are  $116.1$  (2) and  $116.7$  (1)°. The crystal structure is stabilized by  $\text{N}-\text{H}\cdots\text{O}$ ,  $\text{O}-\text{H}\cdots\text{N}$  and  $\text{O}-\text{H}\cdots\text{O}$  hydrogen bonds.

## Related literature

For the uses of organic pyridinium salts, see: Damiano *et al.* (2007). For zwitterionic forms of *N*-arylbenzenesulfonamides, see: Li *et al.* (2007); Yu & Li (2007). For reference geometric data, see: Allen *et al.* (1987).



## Experimental

## Crystal data

$\text{C}_{11}\text{H}_9\text{N}_3\text{O}_4\text{S}\cdot\text{H}_2\text{O}$   
 $M_r = 297.29$   
 Monoclinic,  $P2_1$

$a = 8.7206$  (17) Å  
 $b = 11.972$  (2) Å  
 $c = 12.743$  (3) Å

$\beta = 107.65$  (3)°  
 $V = 1267.8$  (4) Å<sup>3</sup>  
 $Z = 4$   
 Mo  $K\alpha$  radiation

$\mu = 0.28$  mm<sup>-1</sup>  
 $T = 113$  (2) K  
 $0.14 \times 0.10 \times 0.04$  mm

## Data collection

Rigaku Saturn diffractometer  
 Absorption correction: multi-scan  
 (*CrystalClear*; Rigaku/MS, 2005)  
 $T_{\min} = 0.962$ ,  $T_{\max} = 0.989$

10348 measured reflections  
 4633 independent reflections  
 3587 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.033$

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.032$   
 $wR(F^2) = 0.090$   
 $S = 1.03$   
 4633 reflections  
 385 parameters  
 9 restraints

H atoms treated by a mixture of independent and constrained refinement  
 $\Delta\rho_{\max} = 0.22$  e Å<sup>-3</sup>  
 $\Delta\rho_{\min} = -0.40$  e Å<sup>-3</sup>  
 Absolute structure: Flack (1983), 1560 Friedel pairs  
 Flack parameter: 0.52 (6)

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{N4}-\text{H4A}\cdots\text{O9}$	0.913 (10)	1.868 (12)	2.764 (3)	167 (3)
$\text{O9}-\text{H9A}\cdots\text{O1}$	0.864 (10)	2.122 (11)	2.967 (3)	166 (3)
$\text{N1}-\text{H1A}\cdots\text{O10}^{\text{i}}$	0.918 (10)	1.843 (11)	2.757 (3)	174 (4)
$\text{O9}-\text{H9B}\cdots\text{N5}^{\text{ii}}$	0.867 (10)	2.043 (11)	2.901 (3)	170 (2)
$\text{O9}-\text{H9B}\cdots\text{O6}^{\text{ii}}$	0.867 (10)	2.56 (2)	3.125 (3)	123 (2)
$\text{O10}-\text{H10A}\cdots\text{O6}^{\text{iii}}$	0.867 (10)	2.051 (11)	2.914 (3)	174 (4)
$\text{O10}-\text{H10B}\cdots\text{N2}^{\text{iv}}$	0.870 (10)	2.042 (12)	2.902 (3)	170 (3)

Symmetry codes: (i)  $-x + 2, y - \frac{1}{2}, -z + 1$ ; (ii)  $x - 1, y, z$ ; (iii)  $-x + 2, y + \frac{1}{2}, -z + 1$ ; (iv)  $-x + 1, y + \frac{1}{2}, -z + 1$ .

Data collection: *CrystalClear* (Rigaku/MS, 2005); cell refinement: *CrystalClear*; data reduction: *CrystalClear*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HG2431).

## References

- Allen, F. H., Kennard, O., Watson, D. G., Brammer, L., Orpen, A. G. & Taylor, R. (1987). *J. Chem. Soc. Perkin Trans. 2*, pp. S1–19.  
 Damiano, T., Morton, D. & Nelson, A. (2007). *Org. Biomol. Chem.* **5**, 2735–2752.  
 Flack, H. D. (1983). *Acta Cryst.* **A39**, 876–881.  
 Li, J.-S., Chen, L.-G., Zhang, Y.-Y., Xu, Y.-J., Deng, Y., Zeng, T. & Huang, P.-M. (2007). *J. Chem. Res.* pp. 350–352.  
 Rigaku/MS (2005). *CrystalClear*. Rigaku/MS, The Woodlands, Texas, USA.  
 Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.  
 Yu, H.-J. & Li, J.-S. (2007). *Acta Cryst.* **E63**, o3399.

**supplementary materials**

*Acta Cryst.* (2008). E64, o2277 [ doi:10.1107/S1600536808035526 ]

## 2-Nitro-*N*-(4-pyridinio)benzenesulfonamidate monohydrate

L.-B. Hu, J. Chen and C.-Z. Liu

### Comment

Organic pyridinium salts have been widely used in the construction of supramolecular architectures (Damiano *et al.*, 2007). As part of our ongoing studies of supramolecular chemistry involving the pyridinium rings (Li *et al.*, 2007), the structure of the title compound was determined by X-ray diffraction. There are two independent molecules and two independent water molecules in the unit cell. In the cations of the title compound the short C—N distance [N2—C1 = 1.367 (3) Å and N5—C12 = 1.365 (3) Å] indicate the slight conjugation of the sulfonamide N with the pyridinium ring (Allen *et al.*, 1987).

The dihedral angles between the benzene ring and the pyridinium ring are 109.7 (1)° and 110.7 (1)° respectively. And the dihedral angles between the nitro group and the benzene ring are 116.1 (2)° and 116.7 (1)° respectively. The crystal structure is stabilized by N—H···O hydrogen bonds.

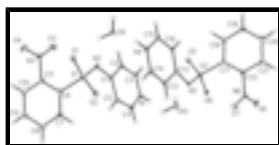
### Experimental

A solution of 2-nitrobenzenesulfonyl chloride (2.2 g, 10 mmol) in CH<sub>2</sub>Cl<sub>2</sub> (10 ml) was added dropwise to a suspension of 4-aminopyridine (0.9 g, 10 mmol) in CH<sub>2</sub>Cl<sub>2</sub> (10 ml) at room temperature with stirring. The reaction mixture was stirred overnight. The yellow solid obtained was washed with warm water to obtain the title compound in a yield of 55.7%. A colorless single crystal suitable for X-ray analysis was obtained by slow evaporation of an NaOH (10%) solution at room temperature over a period of a week.

### Refinement

The N-bound H atoms were located in a difference map and their coordinates were refined with  $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{N})$ . The C-bound H atoms were positioned geometrically (C—H = 0.93 Å) and refined as riding with  $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$ . The Flack test results are ambiguous because of the presence of merohedral twin. The water O-bound H atoms were refined freely, but the O—H distances were restrained to 0.85 (1) Å, and the water H<sub>A</sub>···H<sub>B</sub> distance to 1.45 (1) Å.

### Figures



## 2-Nitro-*N*-(4-pyridinio)benzenesulfonamidate monohydrate

### Crystal data

$C_{11}H_9N_3O_4S \cdot H_2O$	$F_{000} = 616$
$M_r = 297.29$	$D_x = 1.557 \text{ Mg m}^{-3}$
Monoclinic, $P2_1$	Mo $K\alpha$ radiation
Hall symbol: P 2yb	$\lambda = 0.71073 \text{ \AA}$
$a = 8.7206 (17) \text{ \AA}$	Cell parameters from 4439 reflections
$b = 11.972 (2) \text{ \AA}$	$\theta = 2.4\text{--}27.5^\circ$
$c = 12.743 (3) \text{ \AA}$	$\mu = 0.28 \text{ mm}^{-1}$
$\beta = 107.65 (3)^\circ$	$T = 113 (2) \text{ K}$
$V = 1267.8 (4) \text{ \AA}^3$	Block, colourless
$Z = 4$	$0.14 \times 0.10 \times 0.04 \text{ mm}$

### Data collection

Rigaku Saturn diffractometer	4633 independent reflections
Radiation source: rotating anode	3587 reflections with $I > 2\sigma(I)$
Monochromator: confocal	$R_{\text{int}} = 0.033$
$T = 113(2) \text{ K}$	$\theta_{\text{max}} = 27.5^\circ$
$\omega$ scans	$\theta_{\text{min}} = 1.7^\circ$
Absorption correction: multi-scan (CrystalClear; Rigaku/MS, 2005)	$h = -11 \rightarrow 11$
$T_{\text{min}} = 0.962, T_{\text{max}} = 0.989$	$k = -12 \rightarrow 15$
10348 measured reflections	$l = -16 \rightarrow 16$

### Refinement

Refinement on $F^2$	Hydrogen site location: inferred from neighbouring sites
Least-squares matrix: full	H atoms treated by a mixture of independent and constrained refinement
$R[F^2 > 2\sigma(F^2)] = 0.032$	$w = 1/[\sigma^2(F_o^2) + (0.0539P)^2]$
$wR(F^2) = 0.090$	where $P = (F_o^2 + 2F_c^2)/3$
$S = 1.03$	$(\Delta/\sigma)_{\text{max}} = 0.002$
4633 reflections	$\Delta\rho_{\text{max}} = 0.22 \text{ e \AA}^{-3}$
385 parameters	$\Delta\rho_{\text{min}} = -0.40 \text{ e \AA}^{-3}$
9 restraints	Extinction correction: none
Primary atom site location: structure-invariant direct methods	Absolute structure: Flack (1983), with how many Friedel pairs?
Secondary atom site location: difference Fourier map	Flack parameter: 0.52 (6)

### Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) etc. and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

### Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
S1	0.54299 (7)	0.49945 (5)	0.75574 (5)	0.01423 (14)
S2	1.01577 (7)	0.25356 (5)	0.35546 (5)	0.01307 (14)
O1	0.36887 (19)	0.49815 (17)	0.72910 (14)	0.0177 (4)
O2	0.6116 (2)	0.59540 (15)	0.71745 (16)	0.0187 (4)
O3	0.4219 (3)	0.29003 (18)	0.87441 (18)	0.0318 (5)
O4	0.2777 (2)	0.4008 (2)	0.94108 (18)	0.0326 (5)
O5	0.9524 (2)	0.15286 (15)	0.38857 (15)	0.0165 (4)
O6	1.1904 (2)	0.26159 (18)	0.38501 (16)	0.0182 (4)
O7	1.1373 (3)	0.46983 (18)	0.24125 (19)	0.0334 (5)
O8	1.2807 (2)	0.3638 (2)	0.16880 (18)	0.0371 (6)
N1	1.0324 (3)	0.2999 (2)	0.67719 (19)	0.0185 (5)
N2	0.5920 (2)	0.38299 (18)	0.71702 (19)	0.0146 (5)
N3	0.4050 (3)	0.3727 (2)	0.9245 (2)	0.0217 (5)
N4	0.5077 (3)	0.4240 (2)	0.43389 (18)	0.0179 (5)
N5	0.9564 (2)	0.36491 (18)	0.39696 (18)	0.0141 (5)
N6	1.1564 (3)	0.3888 (2)	0.18846 (19)	0.0230 (6)
C1	0.7414 (3)	0.3615 (2)	0.7078 (2)	0.0132 (5)
C2	0.7644 (3)	0.2549 (2)	0.6690 (2)	0.0167 (5)
H2	0.6813	0.2029	0.6537	0.020*
C3	0.9100 (3)	0.2265 (2)	0.6532 (2)	0.0188 (6)
H3	0.9231	0.1563	0.6259	0.023*
C4	1.0168 (3)	0.4019 (2)	0.7153 (2)	0.0179 (6)
H4	1.1030	0.4513	0.7304	0.021*
C5	0.8752 (3)	0.4350 (2)	0.7325 (2)	0.0169 (6)
H5	0.8672	0.5058	0.7604	0.020*
C6	0.6158 (3)	0.5053 (3)	0.9027 (2)	0.0148 (5)
C7	0.7441 (3)	0.5753 (2)	0.9531 (2)	0.0189 (6)
H7	0.7936	0.6156	0.9100	0.023*
C8	0.7996 (3)	0.5858 (2)	1.0674 (2)	0.0240 (6)
H8	0.8857	0.6329	1.0999	0.029*
C9	0.7271 (3)	0.5264 (3)	1.1326 (2)	0.0245 (7)
H9	0.7656	0.5325	1.2089	0.029*
C10	0.5978 (3)	0.4582 (2)	1.0843 (2)	0.0208 (6)

## supplementary materials

H10	0.5469	0.4197	1.1277	0.025*
C11	0.5441 (3)	0.4472 (2)	0.9712 (2)	0.0162 (6)
C12	0.8049 (3)	0.3782 (2)	0.4049 (2)	0.0120 (5)
C13	0.7734 (3)	0.4806 (2)	0.4510 (2)	0.0163 (6)
H13	0.8534	0.5347	0.4722	0.020*
C14	0.6264 (3)	0.4998 (2)	0.4642 (2)	0.0181 (5)
H14	0.6078	0.5669	0.4949	0.022*
C15	0.5302 (3)	0.3265 (2)	0.3884 (2)	0.0165 (5)
H15	0.4464	0.2752	0.3672	0.020*
C16	0.6767 (3)	0.3014 (2)	0.3725 (2)	0.0136 (5)
H16	0.6905	0.2339	0.3405	0.016*
C17	0.9459 (3)	0.2526 (2)	0.2084 (2)	0.0145 (5)
C18	0.8167 (3)	0.1850 (2)	0.1552 (2)	0.0179 (6)
H18	0.7672	0.1423	0.1967	0.021*
C19	0.7603 (3)	0.1799 (2)	0.0417 (2)	0.0229 (6)
H19	0.6723	0.1349	0.0080	0.027*
C20	0.8324 (3)	0.2406 (3)	−0.0223 (2)	0.0245 (7)
H20	0.7933	0.2365	−0.0987	0.029*
C21	0.9631 (3)	0.3074 (2)	0.0276 (2)	0.0195 (6)
H21	1.0141	0.3476	−0.0146	0.023*
C22	1.0173 (3)	0.3137 (2)	0.1415 (2)	0.0157 (6)
O9	0.2408 (2)	0.50194 (18)	0.48520 (16)	0.0183 (4)
O10	0.6923 (2)	0.74683 (18)	0.37557 (18)	0.0203 (4)
H1A	1.127 (3)	0.280 (3)	0.665 (3)	0.054 (11)*
H4A	0.411 (2)	0.441 (2)	0.444 (2)	0.021 (8)*
H9A	0.260 (3)	0.500 (3)	0.5558 (8)	0.047 (11)*
H9B	0.160 (2)	0.460 (2)	0.4517 (18)	0.024 (9)*
H10A	0.719 (4)	0.751 (4)	0.4468 (8)	0.11 (2)*
H10B	0.601 (2)	0.780 (3)	0.346 (2)	0.034 (10)*

### Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
S1	0.0127 (3)	0.0152 (3)	0.0156 (3)	0.0016 (3)	0.0055 (2)	0.0026 (3)
S2	0.0098 (3)	0.0162 (3)	0.0138 (3)	0.0013 (3)	0.0044 (2)	0.0012 (3)
O1	0.0119 (8)	0.0220 (10)	0.0201 (10)	0.0049 (9)	0.0064 (7)	0.0009 (9)
O2	0.0242 (10)	0.0133 (9)	0.0209 (11)	0.0000 (8)	0.0102 (9)	0.0042 (8)
O3	0.0376 (12)	0.0285 (12)	0.0288 (13)	−0.0115 (10)	0.0095 (10)	−0.0019 (10)
O4	0.0157 (9)	0.0515 (15)	0.0329 (13)	−0.0051 (10)	0.0109 (9)	0.0064 (12)
O5	0.0194 (10)	0.0153 (9)	0.0158 (10)	0.0043 (8)	0.0067 (8)	0.0053 (8)
O6	0.0099 (8)	0.0269 (11)	0.0190 (10)	0.0008 (9)	0.0061 (7)	0.0016 (9)
O7	0.0378 (13)	0.0292 (13)	0.0329 (13)	−0.0111 (10)	0.0101 (11)	−0.0060 (10)
O8	0.0180 (10)	0.0689 (18)	0.0274 (13)	−0.0077 (11)	0.0114 (10)	0.0055 (12)
N1	0.0135 (11)	0.0249 (12)	0.0184 (12)	0.0064 (10)	0.0067 (9)	0.0022 (10)
N2	0.0116 (10)	0.0147 (11)	0.0196 (12)	0.0000 (9)	0.0078 (9)	−0.0001 (9)
N3	0.0209 (12)	0.0263 (14)	0.0176 (13)	−0.0040 (11)	0.0057 (10)	0.0071 (10)
N4	0.0148 (11)	0.0224 (13)	0.0179 (12)	0.0020 (10)	0.0072 (9)	0.0009 (10)
N5	0.0108 (10)	0.0177 (11)	0.0144 (11)	−0.0028 (9)	0.0046 (9)	−0.0035 (9)

N6	0.0193 (12)	0.0354 (15)	0.0146 (12)	−0.0067 (11)	0.0054 (10)	0.0064 (11)
C1	0.0123 (11)	0.0178 (13)	0.0100 (13)	0.0024 (10)	0.0042 (10)	0.0041 (10)
C2	0.0166 (11)	0.0179 (12)	0.0164 (13)	−0.0044 (12)	0.0062 (10)	0.0015 (12)
C3	0.0177 (13)	0.0173 (14)	0.0217 (15)	0.0033 (11)	0.0062 (11)	−0.0025 (11)
C4	0.0133 (12)	0.0214 (14)	0.0181 (14)	−0.0009 (11)	0.0035 (10)	0.0014 (12)
C5	0.0145 (12)	0.0147 (13)	0.0211 (15)	0.0009 (11)	0.0048 (11)	−0.0006 (11)
C6	0.0151 (12)	0.0134 (13)	0.0166 (13)	0.0032 (12)	0.0059 (10)	0.0000 (12)
C7	0.0206 (13)	0.0143 (13)	0.0224 (15)	−0.0036 (11)	0.0075 (12)	0.0006 (11)
C8	0.0252 (15)	0.0225 (15)	0.0224 (16)	−0.0067 (13)	0.0040 (13)	−0.0035 (12)
C9	0.0255 (14)	0.0278 (16)	0.0184 (15)	0.0041 (13)	0.0038 (12)	−0.0030 (12)
C10	0.0236 (14)	0.0239 (15)	0.0178 (15)	0.0041 (12)	0.0106 (12)	0.0047 (12)
C11	0.0103 (11)	0.0175 (13)	0.0217 (15)	0.0013 (10)	0.0062 (11)	0.0027 (11)
C12	0.0152 (12)	0.0125 (12)	0.0083 (12)	0.0024 (10)	0.0038 (10)	0.0017 (10)
C13	0.0168 (12)	0.0141 (13)	0.0183 (14)	−0.0048 (11)	0.0060 (11)	−0.0026 (11)
C14	0.0225 (13)	0.0147 (12)	0.0186 (13)	−0.0018 (12)	0.0087 (11)	−0.0020 (12)
C15	0.0136 (12)	0.0198 (13)	0.0173 (14)	0.0002 (11)	0.0066 (11)	−0.0031 (11)
C16	0.0093 (11)	0.0150 (12)	0.0173 (14)	−0.0026 (10)	0.0051 (10)	−0.0032 (11)
C17	0.0160 (12)	0.0150 (12)	0.0119 (13)	0.0007 (12)	0.0033 (10)	−0.0004 (11)
C18	0.0193 (13)	0.0163 (14)	0.0198 (15)	−0.0015 (12)	0.0084 (12)	0.0037 (12)
C19	0.0262 (15)	0.0190 (15)	0.0202 (15)	−0.0061 (13)	0.0022 (12)	0.0005 (12)
C20	0.0322 (15)	0.0275 (16)	0.0115 (14)	0.0032 (14)	0.0033 (12)	−0.0006 (13)
C21	0.0193 (13)	0.0221 (15)	0.0205 (15)	0.0031 (12)	0.0110 (12)	0.0049 (12)
C22	0.0118 (11)	0.0181 (14)	0.0172 (14)	0.0001 (11)	0.0045 (11)	0.0019 (11)
O9	0.0157 (9)	0.0210 (10)	0.0194 (11)	0.0003 (9)	0.0072 (8)	−0.0056 (10)
O10	0.0170 (9)	0.0221 (10)	0.0215 (11)	0.0022 (9)	0.0057 (8)	0.0043 (10)

*Geometric parameters (Å, °)*

S1—O2	1.4464 (19)	C6—C11	1.403 (3)
S1—O1	1.4521 (17)	C7—C8	1.394 (4)
S1—N2	1.581 (2)	C7—H7	0.9300
S1—C6	1.787 (3)	C8—C9	1.383 (4)
S2—O5	1.4416 (19)	C8—H8	0.9300
S2—O6	1.4568 (17)	C9—C10	1.377 (4)
S2—N5	1.578 (2)	C9—H9	0.9300
S2—C17	1.787 (3)	C10—C11	1.380 (4)
O3—N3	1.211 (3)	C10—H10	0.9300
O4—N3	1.237 (3)	C12—C16	1.409 (3)
O7—N6	1.221 (3)	C12—C13	1.421 (3)
O8—N6	1.221 (3)	C13—C14	1.362 (3)
N1—C4	1.336 (4)	C13—H13	0.9300
N1—C3	1.345 (3)	C14—H14	0.9300
N1—H1A	0.918 (10)	C15—C16	1.385 (3)
N2—C1	1.367 (3)	C15—H15	0.9300
N3—C11	1.478 (3)	C16—H16	0.9300
N4—C14	1.342 (3)	C17—C18	1.386 (4)
N4—C15	1.344 (3)	C17—C22	1.404 (3)
N4—H4A	0.913 (10)	C18—C19	1.380 (4)
N5—C12	1.365 (3)	C18—H18	0.9300

## supplementary materials

---

N6—C22	1.482 (3)	C19—C20	1.379 (4)
C1—C2	1.405 (4)	C19—H19	0.9300
C1—C5	1.419 (4)	C20—C21	1.380 (4)
C2—C3	1.387 (3)	C20—H20	0.9300
C2—H2	0.9300	C21—C22	1.385 (4)
C3—H3	0.9300	C21—H21	0.9300
C4—C5	1.376 (3)	O9—H9A	0.864 (10)
C4—H4	0.9300	O9—H9B	0.867 (10)
C5—H5	0.9300	O10—H10A	0.867 (10)
C6—C7	1.389 (4)	O10—H10B	0.870 (10)
O2—S1—O1	116.08 (11)	C9—C8—C7	120.3 (3)
O2—S1—N2	114.62 (10)	C9—C8—H8	119.9
O1—S1—N2	106.27 (11)	C7—C8—H8	119.9
O2—S1—C6	105.68 (13)	C10—C9—C8	119.8 (3)
O1—S1—C6	105.03 (11)	C10—C9—H9	120.1
N2—S1—C6	108.55 (13)	C8—C9—H9	120.1
O5—S2—O6	116.29 (11)	C9—C10—C11	119.8 (3)
O5—S2—N5	114.48 (10)	C9—C10—H10	120.1
O6—S2—N5	106.33 (12)	C11—C10—H10	120.1
O5—S2—C17	105.23 (12)	C10—C11—C6	121.9 (3)
O6—S2—C17	105.61 (11)	C10—C11—N3	117.1 (2)
N5—S2—C17	108.33 (12)	C6—C11—N3	121.0 (2)
C4—N1—C3	121.3 (2)	N5—C12—C16	127.0 (2)
C4—N1—H1A	120 (2)	N5—C12—C13	116.4 (2)
C3—N1—H1A	119 (2)	C16—C12—C13	116.6 (2)
C1—N2—S1	123.21 (19)	C14—C13—C12	120.2 (2)
O3—N3—O4	124.9 (3)	C14—C13—H13	119.9
O3—N3—C11	119.3 (2)	C12—C13—H13	119.9
O4—N3—C11	115.9 (2)	N4—C14—C13	121.6 (3)
C14—N4—C15	120.7 (2)	N4—C14—H14	119.2
C14—N4—H4A	118.3 (19)	C13—C14—H14	119.2
C15—N4—H4A	121.0 (19)	N4—C15—C16	120.9 (2)
C12—N5—S2	123.16 (18)	N4—C15—H15	119.6
O7—N6—O8	125.8 (3)	C16—C15—H15	119.6
O7—N6—C22	118.2 (2)	C15—C16—C12	120.0 (2)
O8—N6—C22	115.9 (2)	C15—C16—H16	120.0
N2—C1—C2	116.3 (2)	C12—C16—H16	120.0
N2—C1—C5	127.2 (2)	C18—C17—C22	116.9 (2)
C2—C1—C5	116.5 (2)	C18—C17—S2	119.0 (2)
C3—C2—C1	120.6 (2)	C22—C17—S2	124.1 (2)
C3—C2—H2	119.7	C19—C18—C17	121.1 (2)
C1—C2—H2	119.7	C19—C18—H18	119.4
N1—C3—C2	120.3 (2)	C17—C18—H18	119.4
N1—C3—H3	119.8	C20—C19—C18	121.0 (3)
C2—C3—H3	119.8	C20—C19—H19	119.5
N1—C4—C5	121.1 (3)	C18—C19—H19	119.5
N1—C4—H4	119.5	C19—C20—C21	119.6 (3)
C5—C4—H4	119.5	C19—C20—H20	120.2
C4—C5—C1	120.2 (2)	C21—C20—H20	120.2



C4—C5—H5	119.9	C20—C21—C22	119.1 (3)
C1—C5—H5	119.9	C20—C21—H21	120.4
C7—C6—C11	117.3 (2)	C22—C21—H21	120.4
C7—C6—S1	119.3 (2)	C21—C22—C17	122.3 (2)
C11—C6—S1	123.3 (2)	C21—C22—N6	115.7 (2)
C6—C7—C8	120.9 (3)	C17—C22—N6	122.0 (2)
C6—C7—H7	119.6	H9A—O9—H9B	111.6 (16)
C8—C7—H7	119.6	H10A—O10—H10B	110.7 (16)
O2—S1—N2—C1	36.7 (2)	O4—N3—C11—C10	62.9 (3)
O1—S1—N2—C1	166.3 (2)	O3—N3—C11—C6	64.4 (4)
C6—S1—N2—C1	-81.2 (2)	O4—N3—C11—C6	-116.0 (3)
O5—S2—N5—C12	-36.0 (2)	S2—N5—C12—C16	-4.2 (4)
O6—S2—N5—C12	-165.8 (2)	S2—N5—C12—C13	175.7 (2)
C17—S2—N5—C12	81.0 (2)	N5—C12—C13—C14	-178.5 (2)
S1—N2—C1—C2	-178.07 (19)	C16—C12—C13—C14	1.4 (4)
S1—N2—C1—C5	1.5 (4)	C15—N4—C14—C13	-0.6 (4)
N2—C1—C2—C3	177.8 (2)	C12—C13—C14—N4	-0.4 (4)
C5—C1—C2—C3	-1.9 (4)	C14—N4—C15—C16	0.6 (4)
C4—N1—C3—C2	-0.8 (4)	N4—C15—C16—C12	0.4 (4)
C1—C2—C3—N1	1.4 (4)	N5—C12—C16—C15	178.5 (2)
C3—N1—C4—C5	0.7 (4)	C13—C12—C16—C15	-1.3 (4)
N1—C4—C5—C1	-1.2 (4)	O5—S2—C17—C18	18.6 (2)
N2—C1—C5—C4	-177.8 (3)	O6—S2—C17—C18	142.2 (2)
C2—C1—C5—C4	1.8 (4)	N5—S2—C17—C18	-104.2 (2)
O2—S1—C6—C7	-14.8 (2)	O5—S2—C17—C22	-159.0 (2)
O1—S1—C6—C7	-138.0 (2)	O6—S2—C17—C22	-35.5 (3)
N2—S1—C6—C7	108.6 (2)	N5—S2—C17—C22	78.1 (2)
O2—S1—C6—C11	161.6 (2)	C22—C17—C18—C19	-1.0 (4)
O1—S1—C6—C11	38.3 (3)	S2—C17—C18—C19	-178.8 (2)
N2—S1—C6—C11	-75.0 (2)	C17—C18—C19—C20	1.1 (4)
C11—C6—C7—C8	0.6 (4)	C18—C19—C20—C21	0.1 (4)
S1—C6—C7—C8	177.2 (2)	C19—C20—C21—C22	-1.2 (4)
C6—C7—C8—C9	-0.1 (4)	C20—C21—C22—C17	1.3 (4)
C7—C8—C9—C10	-1.1 (4)	C20—C21—C22—N6	-179.0 (2)
C8—C9—C10—C11	1.7 (4)	C18—C17—C22—C21	-0.2 (4)
C9—C10—C11—C6	-1.2 (4)	S2—C17—C22—C21	177.5 (2)
C9—C10—C11—N3	180.0 (2)	C18—C17—C22—N6	-179.9 (2)
C7—C6—C11—C10	0.0 (4)	S2—C17—C22—N6	-2.2 (4)
S1—C6—C11—C10	-176.4 (2)	O7—N6—C22—C21	116.6 (3)
C7—C6—C11—N3	178.8 (2)	O8—N6—C22—C21	-62.4 (3)
S1—C6—C11—N3	2.4 (4)	O7—N6—C22—C17	-63.7 (4)
O3—N3—C11—C10	-116.7 (3)	O8—N6—C22—C17	117.3 (3)

Hydrogen-bond geometry (Å, °)

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
N4—H4A $\cdots$ O9	0.913 (10)	1.868 (12)	2.764 (3)	167 (3)
O9—H9A $\cdots$ O1	0.864 (10)	2.122 (11)	2.967 (3)	166 (3)
N1—H1A $\cdots$ O10 <sup>i</sup>	0.918 (10)	1.843 (11)	2.757 (3)	174 (4)

## supplementary materials

---

O9—H9B···N5 <sup>ii</sup>	0.867 (10)	2.043 (11)	2.901 (3)	170 (2)
O9—H9B···O6 <sup>ii</sup>	0.867 (10)	2.56 (2)	3.125 (3)	123 (2)
O10—H10A···O6 <sup>iii</sup>	0.867 (10)	2.051 (11)	2.914 (3)	174 (4)
O10—H10B···N2 <sup>iv</sup>	0.870 (10)	2.042 (12)	2.902 (3)	170 (3)

Symmetry codes: (i)  $-x+2, y-1/2, -z+1$ ; (ii)  $x-1, y, z$ ; (iii)  $-x+2, y+1/2, -z+1$ ; (iv)  $-x+1, y+1/2, -z+1$ .

Fig. 1

